

## AC E 506 : ANALYTICAL AND GREEN CHEMISTRY

### COURSE OUTCOME:

Enable the students:

- To understand the basic principles and theory of UV-Visible, Electronic, Infra Red, Nuclear Magnetic Resonance and Mass Spectroscopy.
- To study the utility of these techniques in structure elucidation of simple organic molecules.
- To know about water cycle, water sources, water quality, significant measurements of water parameters and treatment of water for drinking and industrial purposes.
- To learn about principles and use of green chemistry in laboratory synthesis.
- To understand the basic principles and utility of sonochemistry and Microwave induced organic synthesis.

### UNIT- I:

[12 Hours]

**UV/Electronic Spectroscopy:** Basic principles, Beer-Lambert law, types of absorption bands, Factors affecting the positions of UV bands. Theoretical prediction of  $\lambda_{\text{max}}$  for polyenes,  $\alpha, \beta$ -unsaturated aldehydes, ketones (Woodward-Fieser rules) and substituted benzenes.

**IR Spectroscopy:** Basic principles, Application of infrared spectroscopy in the structural study-identity by finger printing and identification of functional groups. Characteristic vibrational frequencies of alkanes, alkenes, alkynes, aromatic compounds, alcohols, ethers, phenols and amines). Study of vibrational frequencies of carbonyl compounds (ketones, aldehydes, esters, amides and acids). Factors affecting band positions and intensities

**Nuclear Magnetic Resonance Spectroscopy:** Basic principles, Solvents used, chemical shift and its measurements, factors affecting chemical shift. Integration of NMR signals, spin-spin coupling, coupling constant. Shielding and deshielding. High resolution  $^1\text{H}$  NMR. Applications of NMR spectroscopy in structure elucidation of simple organic molecules.

**Mass Spectrometry:** Basic principles, molecular ions, meta-stable ions and isotope ions. Fragmentation processes, McLafferty rearrangement. retro Diels-Alder fragmentations. Nitrogen rule.

### UNIT- II:

[12 Hours]

**Hydrologic cycle, sources, chemistry of sea water, criteria and standards of water quality- safe drinking water, maximum contamination levels of inorganic and organic chemicals, radiological contaminants, turbidity, microbial contaminants. Public health significance and measurement of colour, turbidity, total solids, acidity, alkalinity, hardness, chloride, residual chlorine, sulphate, fluoride, phosphate and different forms of nitrogen in natural and polluted water. Chemical sources of taste and odour, treatment for their removal, sampling and monitoring techniques. Determination and significance of DO, BOD, COD and TOC. Water purification for drinking and industrial purposes, disinfection techniques, demineralization, desalination processes and reverse osmosis. Treatment of liquid radioactive wastes**

### UNIT- III:

[12 Hours]

**Green Chemistry:** Definition and principles, planning a green synthesis in a chemical laboratory, Green preparation-Aqueous phase reactions, solid state (solventless) reactions, photochemical reactions, Phase transfer catalyst catalysed reactions (Quaternary ammonium salts & Crown ethers), enzymatic transformations & reactions in ionic liquids.

**Sonochemistry:** Introduction, instrumentation, the phenomenon of cavitation, Sonochemicalesterification, substitution, addition, oxidation, reduction and coupling reactions.

**Microwave induced organic synthesis:** Introduction, reaction vessel and reaction medium, concept, specific effect, atom efficiency, % atom utilisation, advantages and limitations, alkylation of active methylene compounds, N-alkylation, condensation of active methylene compounds with aldehydes, Diels-Alder reaction, Leuckardt reductive amination of ketones, ortho ester Claisen rearrangement.

#### References:-

1. Organic Spectroscopy-3<sup>rd</sup> Ed.-W.Kemp (Pargrave Publishers, New York), 1991.
2. Spectrometric Identification of Organic Compounds - Silverstein, Bassler&Monnill (Wiley)1981.
3. Applications of Absorption Spectroscopy of Organic Compounds-Dyer(Prentice Hall,NY) 1965.
4. Spectroscopy of Organic Compounds-3<sup>rd</sup> Ed.-P.S.Kalsi (New Age, New Delhi) 2000.
5. Spectroscopic Methods in Organic Chemistry - Williams and Fleming, TMH.
6. A.K. De : Environmental Chemistry, (Wiley Eastern).
7. S.K.Banerji : Environmental Chemistry, ( Prentice Hall India), 1993.
- 8 S.D. Faust and O.M. Aly : Chemistry of Water Treatment, (Butterworths),1983.
9. Sawyer and McCarty, Chemistry for Environmental Engineering(McGraw Hill) 1978
10. I.Williams, Environmental Chemistry, John Wiley, 2001.
- S.M.Khopkar, Environmental Pollution Analysis, (Wiley Eastern).
10. Organic Synthesis-Special Techniques, V.K.Ahluwalia& R. Aggarwal, Narosa, 2001.
11. Green Chemistry-Environment friendly alternatives- R.Sanghi&M.M.Srivatsava, Narosa, 2003. 14. Green Chemistry-Environment benign reactions- V.K.Ahluwalia, Ane Books India, 2006.

### AC P 507: INORGANIC CHEMISTRY PRACTICALS

#### COURSE OUTCOME:

- The students will have hands on experience in the Analysis of Brass, Cu-Ni alloy, Stainless Steel,
- Type Metal and quantitative analysis of the constituents & mixtures containing the following radicals Fe + Ni, Fe + Ca, Cr + Fe.
- This course also train the students in Separation and determination of  $Mg^{2+}/Zn^{2+}$ ,  $Zn^{2+}/Cd^{2+}$  by Ion-Exchange Chromatography and ion exchange capacity of a resin.

1. Analysis of brass—Cu gravimetrically using  $\alpha$ -Benzoinoxime& Zn complexometrically.
2. Analysis Cu-Ni alloy .
3. Analysis of Stainless Steel – Insoluble residue by gravimetry, Ni gravimetrically using DMG, Fe volumetrically using Ce(IV) & Cr volumetrically by persulphate oxidation.
4. Analysis of Type metal –Sn gravimetrically, Pb electrogravimetrically and Sb titrimetrically using  $KBrO_3$
5. Quantitative analysis of the constituents & mixtures containing the following radicals.  
Cu(II) + Fe(II) - Cu gravimetrically as  $CuSCN$  and Fe using Ce(IV).  
Fe(II) + Ni(II) – Fe gravimetrically as  $Fe_2O_3$  and Ni using EDTA.
8. Fe(III) + Ca(II) - Fe gravimetrically as  $Fe_2O_3$  and Ca using EDTA.
9. Cr(III) + Fe(III) – Using EDTA by Kinetic masking method.
6. Analysis of chalcopyrites, magnetite and ilmenite.
7. Ion-exchange chromatography: Separation and determination of  $Mg^{2+} / Zn^{2+}$ ,  $Zn^{2+} / Cd^{2+}$ ;  $Cl^- / Br^-$
8. Separation of cations using column and paper chromatography